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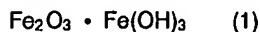
(11) Publication number:

**0 630 856 A1**

(12)

**EUROPEAN PATENT APPLICATION**(21) Application number: **94109023.5**(51) Int. Cl.<sup>5</sup>: **C01G 9/04, C01G 9/00,  
C01G 49/06, C01G 21/00**(22) Date of filing: **13.06.94**(30) Priority: **28.06.93 IT MI931384**(43) Date of publication of application:  
**28.12.94 Bulletin 94/52**(84) Designated Contracting States:  
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**I-20121 Milano (IT)**(54) **Process for removing iron and manganese from acid solutions of zinc salts.**

(57) A process for removing iron and manganese from solutions of zinc chloride containing hydrochloric acid, having a zinc ion concentration between 50 and 500 gm/l, an iron ions concentration between 1 and 70 gm/l, and a manganese concentration between 10 and 200 ppm, comprising a neutralization of free acid; a first oxidation step of ferrous ions to ferric ones by means of a first oxidizing agent selected in the group consisting of sodium chlorate, hydrogen peroxide, sodium hypochlorite, sodium chlorite; and an iron precipitation as a compound with formula (1)



in the form of easily filterable crystalline particles. After the beginning of the crystallization a second oxidizing agent formed by potassium permanganate is added, so that manganese is precipitated as  $\text{MnO}_2$ ; then the compound with formula (1) and  $\text{MnO}_2$  are

removed together by filtering.

**EP 0 630 856 A1**

The present invention relates to a process for removing iron and manganese from acid solutions of zinc salts. Generally the field of the present invention regards polluted solutions obtained in galvanizing plants, in pickling iron or badly plated parts before zinc-plating. Particularly the invention relates to a solution containing hydrochloric acid that has a zinc ion concentration between 50 and 500 gm/l, preferably between 150 and 400 gm/l, and an iron ions concentration between 1 and 70 gm/l, and a manganese concentration between 10 and 200 ppm, preferably between 50 and 150 ppm.

Since zinc solutions containing iron ions, usually ferrous ions, and other impurities, can hardly be recycled to other purposes, their treatment is often very expensive.

On the contrary it is common knowledge that ironless and manganeseless zinc solutions may be easily employed again in other industrial operations, sometimes in the plating plants themselves, and the iron removed from zinc may be recycled to other purposes.

Furthermore zinc solutions free from iron, manganese and heavy metals can be used in more interesting industrial processes with stricter requirements. This is particularly important because in this way the whole process can be remunerative.

It is common knowledge that neutralization of iron solutions, e.g. with sodium hydroxide, involves precipitations of iron hydroxide, which is very difficult to filter; moreover, increasing pH-value to about 6, also zinc precipitates together with iron as  $\text{Zn(OH)}_2$ .

It is also known that the ferrous ions may be oxidized to ferric ions, so that iron precipitates as  $\text{Fe(OH)}_3$ . However applicant found that  $\text{Fe(OH)}_3$  gives rise to a colloidal precipitate and its filtration is difficult to perform, for the  $\text{Fe(OH)}_3$  slime obstructs filters; consequently filtration time is prolonged beyond any industrial benefit. Moreover, the obtained slimes are very gelatinous and retain mother liquor in plenty (they have a high titre of zinc, for instance about 30%).

US Patent No. 4 591 489 and French Patent No. 2 307 880 disclose processes for removing iron from zinc and iron solutions. According to these processes, an organic layer and a distinct aqueous one are used. Many additional plant and maintenance costs, which generally are exorbitant compared with the low value of these solutions, are involved if an organic layer is employed.

US Patent No. 3 600 128 discloses a removal process based on the use of activated carbon. The carbon increases the filtering rate. Nevertheless, the iron is necessarily removed together with carbon and cannot be employed for other utilizations.

Rumanian Patent No 96 826 discloses a process for recovering only a zinc sulfate solution, wherein

a paste of zinc oxide and hydrogen peroxide is proposed. Nevertheless, the preparation of this mixture involves risk of decomposition and also of explosion if metallic impurities are present.

Furthermore none of the above prior art process discloses any means to remove manganese.

Egypt. J. Chem., 15, No 2, 175-178 (1972) discloses a similar process. an oxidizing treatment with  $\text{H}_2\text{O}_2$  and an iron precipitation with  $\text{ZnO}$ . However no means to remove manganese from the solution is disclosed. Furthermore at the end of the process about 21 mg of iron and 104 mg of lead are still present for each kg of zinc. These amounts are by far not acceptable for recycling zinc.

FR 2 042 960 filed on 1970 discloses an oxidizing treatment with sodium chlorate and an iron precipitation with a mixture of  $\text{NH}_4\text{OH}$  and  $\text{ZnO}$ . Also in this case there is no teaching for the removal of manganese. Furthermore thousands of ppm of  $\text{NH}_4^+$  are introduced into the solution and cannot be removed. Finally it has been found that it is very difficult to filter the precipitated product.

DE 1 817 611 filed on 1968 discloses the treatment of wastes containing  $\text{H}_2\text{SO}_4$  and not  $\text{HCl}$  and is therefore outside the field of the present invention. The oxidizing treatment is carried out with air and the precipitation with  $\text{ZnO}$ . Also a treatment with zinc sulfide is described. Also in this case there is no teaching for the manganese removal. In any case, at the end of the process 20000 mg of iron 10000 mg of copper, 1721 mg of cadmium are still present for each kg of zinc.

US 1 973 300 filed on 1932 and SU 1 479 419 teach exclusively the treatment of zinc sulphate solutions and therefore are outside the field of the present invention, also do not disclose any teaching to remove manganese.

Literature provides also electrochemical removal processes, whose complexity involves expensive operations and difficulties in order to obtain a sufficiently pure zinc.

Even if in more than 60 years many attempts were done to remove iron from zinc solutions, nevertheless, in industry zinc and iron solutions are not actually separated and recycled but only treated in purification plants; slime rises, which involves problems for environment and additional expenses.

An aim of the present invention is to overcome said disadvantages and to obtain an industrial process which can be carried out in simple plants, at low cost and without needing specialized workers.

Another aim of the present invention is to remove iron quickly, though filtering with normal equipments.

A further aim is to obtain zinc solutions of high purity and very low iron content.

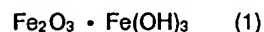
Another aim is to remove manganese which is a critical polluting element and prevents recycling

in many other industrial uses.

A further object is the purification of zinc solutions not only from iron and manganese, but also from other impurities, particularly heavy metals, so that cleared solutions are obtained, suitable e.g. in production of dithiocarbammates, used as agricultural fungicides, or rubber polymerization agents, or in the plating plants themselves mixed with ammonium chloride.

Another object is the possibility of using iron, for instance as pigment in cement or bricks in building industries.

All these aims and other objects are achieved by the process, according to the present invention, for removing iron and manganese from solutions of zinc salts containing hydrochloric acid, having a zinc ion concentration between 50 and 500 gm/l, an iron ions concentration between 1 and 70 gm/l, and a manganese concentration between 10 and 200 ppm, characterized in that it comprises: a neutralization of free acid; a first oxidation step of ferrous ions to ferric ones by means of a first oxidizing agent selected in the group consisting of sodium chlorate, hydrogen peroxide, sodium hypochlorite, sodium chlorite; and an iron precipitation as a compound with formula (1)



in the form of easily filterable crystalline particles, said precipitation being obtained by means of an addition of zinc oxide before or after said first oxidation step; after the beginning of said crystallization a second oxidizing agent formed by potassium permanganate being added, so that manganese is precipitated as  $\text{MnO}_2$ ; then the compound with formula (1) and  $\text{MnO}_2$  being removed together by filtering.

The compound with formula (1) has a crystalline form and, above all, is easily filtered.

Preferably after the neutralization step, an amount of zinc oxide is added as stoichiometrically needed for said iron precipitation; this addition is followed by said oxidation causing a simultaneous iron precipitation.

The neutralization step can be performed with metallic zinc, economically advantageous, even if hydrogen is involved and ventilation is required; zinc oxide is preferably used.

The neutralization step increases pH-value without involving any iron precipitation and it is performed indifferently at room temperature or warming.

According to a possible process, the stoichiometrically required amount of zinc oxide can be added after the oxidation so that iron precipitation occurs after the oxidation. According to a preferred process, zinc oxide is added to the zinc

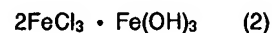
salt solution in such an amount as needed for both neutralization step and iron precipitation step; however precipitation does not occur at this stage. Oxidation step is performed after zinc oxide addition and consequently oxidation and precipitation occurs at the same time.

The possibility of obtaining precipitation of the compounds of formula (1) in the oxidation step itself is very advantageous in practice. In this case the amount of zinc oxide that is added to the solution before oxidation step must be stoichiometrically sufficient to cause precipitation of all the iron present in the solution.

In the preferred process, oxidation step is performed by adding sodium chlorate. In particular, sodium chlorate, compared with other oxidizing substances, as for instance sodium hypochlorite or chlorite, has following advantages: it has a low cost, it adds a restricted amount of sodium ions to solutions, it reacts quickly at relatively low temperatures.

On the other hand, hydrogen peroxide, even if it has the advantage that no sodic ions are added to the solution, is economically disadvantageous.

The oxidation step is performed after neutralization of free acid. Therefore the compound with formula (2) is obtained



from which the compound of formula (1) can be easily produced.

The oxidation step is preferably carried out at temperature between  $20^\circ\text{C}$  and the boiling point of the solution, preferably between  $75^\circ\text{C}$  and  $90^\circ\text{C}$ .

Preferably the second oxidizing agent i.e. potassium permanganate is added after at least 15 minutes and more preferably after 30 minutes from the beginning of the iron precipitation. It has been found that this time period prevents that iron, instead of manganese, is oxidized by the second oxidizing agent.

Preferably during the second oxidation the reaction medium is stirred to improve the contact between the compound of formula (1) and  $\text{MnO}_2$ .

Applicant has found that the precipitation of manganese with potassium permanganate cannot be obtained before iron precipitation or after that iron has been removed from the medium. In fact, before iron precipitation the addition of potassium permanganate may give rise to a partial iron precipitation, but not to manganese precipitation. After that iron has been removed from the medium the addition of potassium permanganate does not give rise to any manganese precipitation at all. In this last case it is possible to obtain only a pale yellow colloidal suspension of  $\text{MnO}_2$  which does not precipitate at all and which cannot be filtered.

Surprisingly  $\text{MnO}_2$  can be easily removed if potassium permanganate is added when the crystals of the compound of formula (1) are still present in the medium. In this last case the obtained suspension easily decants. However it is preferable to filter the solution because this operation does not give rise to problems and is extremely fast. In other words there is a synergistic effect in the process according to claim 1 because the remotion of iron as compound of formula (1) and of manganese as  $\text{MnO}_2$  can be carried out only at the same time, as claimed.

It is preferable to carry out a limited washing of the filtered panel with water to recover more than 95% of zinc chloride impregnating the panel.

It has been noted that the panel contains not only iron and manganese, but also the major part of the polluting elements originally present into the starting solution, because, during the process a coprecipitation of many inorganic impurities occurs. For example copper decreases from a starting concentration higher than 5 ppm to a final concentration of less than 1 ppm, and also lead concentration is substantially reduced.

After the above filtration, a preferred subsequent step of the process is the addition to the solution of a stoichiometrical amount of a soluble alkaline sulphide, preferably ammonium sulphide. This addition gives rise to the precipitation of antimony, arsenic, copper, mercury, tin, bismuth. In fact it is obtained a light precipitate of the sulphides of the above metals. This precipitate can be easily recovered with a common paper or cloth filter. It has been found that this precipitation does not occur if the sulphide addition is carried out before the filtering of the compound of formula (1) and manganese.

It has been found that in the above conditions surprisingly lead sulphide does not precipitate, even changing pH. It should be noted that the treatment of solution having a high saline concentration of zinc chloride is particularly difficult because these solutions have a surprising behaviour.

Preferably a cationic ion exchange resin is used to remove the remaining amount of lead and other little amounts of impurities. Advantageously the remotion of the major part of the polluting elements with the preceding treatments dramatically decreases the charge on the ion exchange resin, so that the life of the resin before regeneration is extremely long.

The invention will be better explained by following examples, whose purpose is descriptive and not restrictive.

#### Example 1

1 liter of polluted starting zinc chloride solution contains: 190 g/l of zinc, 16 g/l of iron, 19 g/l of hydrochloric acid, 100 mg/l of manganese, and about 32 mg/l of lead, 5 mg/l of antimony, 1 mg/l of arsenic, 3 mg/l of copper, 0,01 mg/l of mercury, 1 mg/l of tin, 1 mg/l of bismuth. The solution is set in a stirred Becker. 46.8 g of  $\text{ZnO}$  are added to the solution, i.e. a light excess with respect to the stoichiometric amount. The  $\text{ZnO}$  amount added to the solution is required for acid neutralization and iron precipitation.

A cloud of the solution is noted for the formation of oxichlorides, while pH rises over 5.

The solution is then warmed at about  $85^\circ\text{C}$  and 30.5 cc of a 250 g/l solution of sodium chlorate are added. Precipitation of the compound of formula (1) is obtained.

After 1 hour 200 mg of  $\text{KMnO}_4$  are added under stirring. The solution is filtered in a sintered porous G-3 Buchner funnel, diameter 120 mm, under suction. Time required: 4 minutes. The panel is washed with 70 cc of water. The obtained solution has  $\text{pH} = 4.5$  at  $20^\circ\text{C}$ .

The panel contains very limited amounts of incorporated zinc: about 0.16 g. Therefore the lost of this ion is negligible.

The obtained solution contains: 212 g/l of Zn, less than 1 ppm of iron, 2 mg/l of manganese, 15 mg/l of lead, 4 mg/l of antimony, 1 mg/l of arsenic, 0,5 mg/l of copper, 0,01 mg/l of mercury, 0,8 mg/l of tin, 0,6 mg/l of bismuth.

#### Example 2 (comparison)

Example 1 is repeated except that  $\text{KMnO}_4$  is not added to the solution.

The obtained solution is pale yellow and contains: 212 g/l of Zn, less than 1 ppm of iron, 98 mg/l of manganese, 15 mg/l of lead, 4 mg/l of antimony, 1 mg/l of arsenic, 0,5 mg/l of copper, 0,01 mg/l of mercury, 0,8 mg/l of tin, 0,6 mg/l of bismuth.

#### Example 3 (comparison)

Example 1 is repeated except that  $\text{KMnO}_4$  is added to the solution after the filtering of the compound of formula (1). Then the solution is filtered again.

The obtained solution is pale yellow and contains: 212 g/l of Zn, less than 1 ppm of iron, 72 mg/l of manganese, 16 mg/l of lead, 4 mg/l of antimony, 1 mg/l of arsenic, 0,6 mg/l of copper, 0,01 mg/l of mercury, 0,8 mg/l of tin, 0,6 mg/l of bismuth.

Example 4

1.7 cc of 0.1 mol/l ammonium sulphide are added to the solution obtained at the end of example 1 after filtration, at a temperature of 40°C. A light cloud is obtained and after 1 hour the solution is filtered on paper.

The obtained solution contains: 212 g/l of Zn, less than 1 ppm of iron, 2 mg/l of manganese, 15 mg/l of lead, 1 mg/l of antimony, 0.2 mg/l of arsenic, 0.2 mg/l of copper, 0.001 mg/l of mercury, 0.2 mg/l of tin, 0.2 mg/l of bismuth.

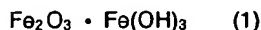
Example 5

The solution obtained at the end of example 4 is finally treated with a previously activated cationic ion exchange resin. The column is 40 cm high and has an apparent volume of about 250 cc. The complete treatment is carried out in about 30 minutes.

The obtained solution contains: 212 g/l of Zn, less than 0.5 ppm of iron, 2 ppm of manganese, 1 mg/l of lead, 0.1 mg/l of antimony, 0.1 mg/l of arsenic, 0.1 mg/l of copper, 0.001 mg/l of mercury, 0.1 mg/l of tin, 0.1 mg/l of bismuth.

Claims

1. A process for removing iron and manganese from solutions of zinc chloride containing hydrochloric acid, having a zinc ion concentration between 50 and 500 gm/l, an iron ions concentration between 1 and 70 gm/l, and a manganese concentration between 10 and 200 ppm, characterized in that it comprises: a neutralization of free acid; a first oxidation step of ferrous ions to ferric ones by means of a first oxidizing agent selected in the group consisting of sodium chlorate, hydrogen peroxide, sodium hypochlorite, sodium chlorite; and an iron precipitation as a compound with formula (1)



in the form of easily filterable crystalline particles, said precipitation being obtained by means of an addition of zinc oxide before or after said first oxidation step; after the beginning of said crystallization a second oxidizing agent formed by potassium permanganate being added, so that manganese is precipitated as  $\text{MnO}_2$ ; then the compound with formula (1) and  $\text{MnO}_2$  being removed together by filtering.

2. A process as in claim 1, wherein after said neutralization step an amount of zinc oxide is added, and then said first oxidation step fol-

lows, while precipitation takes place at the same time.

3. A process according to at least one of the preceding claims, wherein the solutions of zinc salt are treated with an amount of zinc oxide sufficient for both neutralization step and iron precipitation, and this operation is followed by said first oxidation step, so that oxidation and precipitation occur simultaneously.
4. A process according to at least one of the preceding claims, wherein said first oxidation step is obtained with sodium chlorate.
5. A process according to at least one of the preceding claims, wherein the process is performed at a temperature between 20°C and the boiling point of the solution, preferably between 75° and 90°C.
6. A process according to at least one of the preceding claims, wherein said neutralization step is carried out by means of the addition of metallic zinc, zinc oxide, carbonate, bicarbonate, hydroxide of an alkaline or alkaline-earth metal, or mixtures thereof.
7. A process according to at least one of the preceding claims in which zinc oxide is added in a stoichiometric amount with respect to iron.
8. A process according to at least one of the preceding claims in which said second oxidizing agent is added after at least 15 minutes and preferably after 30 minutes from the beginning of said iron precipitation.
9. A process according to at least one of the preceding claims in which the reaction medium is stirred.
10. A process according to at least one of the preceding claims in which after said filtering, heavy metals, except lead, are precipitated by means of addition of a sulfide, preferably ammonium sulfide, then precipitated heavy metals are filtered.
11. A process according to claim 10 comprising a subsequent treatment with a cationic ion exchange resin to remove lead and eventual remaining polluting elements.



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## EUROPEAN SEARCH REPORT

Application Number  
EP 94 10 9023

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.5)
Y,D	FR-A-2 042 960 (EMPRESA AUXILIAR DE LA INDUSTRIA) * the whole document *	1-3,5	C01G9/04 C01G9/00 C01G49/06 C01G21/00
X	---	4	
Y	US-A-3 148 944 (C.P. VAN DIJK) * claim 9 *	1-3,5	
X	---	11	
	DATABASE WPI Derwent Publications Ltd., London, GB; AN 77-35053Y & JP-A-52 042 420 (TOHO BESLON) * abstract *		
X	---	10	
	DATABASE WPI Derwent Publications Ltd., London, GB; AN 80-66078C & DD-A-142 701 (VEB LEUNA-WERK W ULBRICH) * abstract *		
A,D	EGYPTIAN JOURNAL OF CHEMISTRY, vol.15, no.2, 1972, CAIRO pages 175 - 178 A.H. KAMEL ET AL 'preparation of pure zinc sulphate and zinc chloride from galvanization skimmings' * the whole document *	1-7	TECHNICAL FIELDS SEARCHED (Int.Cl.5)  C01G C22B
A,D	---		
	DE-A-18 17 611 (SOC. DES MINES ET FONDERIES DE ZINC DE LA VIEILLE MONTAGNE) * page 3 - page 4 *	1-7	
A,D	---		
	US-A-1 973 300 (J.C. THOMPSON) * the whole document *	1-7	
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The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 5 October 1994	Examiner LIBBERECHT, E
CATEGORY OF CITED DOCUMENTS			
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document		I : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons ----- & : number of the same patent family, corresponding document	